

B.N. Bandodkar College of Science, Thane
S.Y.B.Sc Second Term End Examination MARCH 2012

Chemistry

Paper I

Duration 2 hrs

Max Marks-60

N.B. : 1) All questions are compulsory

2) Figures to the right indicate full marks.

3) Use of log tables/non-programmable calculator is allowed.

- Q. 1 A a) Define the term component with a suitable example. 2
- OR
- a) What is meant by CST ? 2
- b) State Raoult's law. 2
- c) Give an example of one component system. 1
- B a) Define Galvanic cell. 2
- OR
- a) Mention the types of ion specific electrode. 2
- b) Write the individual electrode reaction and the net cell reaction of the following cell 1
- $\text{Ag} \mid \text{Ag}^+ \parallel \text{Fe}^{2+} \mid \text{Fe}$
- OR
- b) How is acidic buffer prepared? 1
- c) Define buffer action. 1
- d) Write the reaction of the following electrode. 1
- $\text{Pt} \mid \text{Fe}^{2+}, \text{Fe}^{3+}$
- C 2
- a) Define titrant and titrand. 2
- OR
- a) State the principle of conductometric titrations. 2
- b) Define titration error. 2
- c) Mention the main groups of titrimetric methods. 1

PTO

